

**ATOMIC AND MOLECULAR PHYSICS SOLUTIONS****GATE-2010**

Q1. To detect trace amounts of gaseous species in a mixture of gases, the preferred probing tool is

- (a) Ionization spectroscopy with  $X$ -rays                      (b) NMR spectroscopy  
(c) ESR spectroscopy    (d) Laser spectroscopy

Ans: (a)

Q2. A collection of  $N$  atoms is exposed to a strong resonant electromagnetic radiation with  $N_g$  atoms in the ground state and  $N_e$  atoms in the excited state, such that  $N_g + N_e = N$ . This collection of two-level atoms will have the following population distribution:

- (a)  $N_g \ll N_e$                       (b)  $N_g \gg N_e$                       (c)  $N_g \approx N_e \approx N/2$                       (d)  $N_g - N_e \approx N/2$

Ans: (c)

Solution: In two level laser population inversion is possible to achieve at any power level. The

maximum possible situation can be  $N_g \approx N_e \approx \frac{N}{2}$

Q3. Two states of an atom have definite parities. An electric dipole transition between these states is

- (a) Allowed if both the states have even parity  
(b) Allowed if both the states have odd parity  
(c) Allowed if the two states have opposite parities  
(d) Not allowed unless a static electric field is applied

Ans: (c)

Q4. The spectrum of radiation emitted by a black body at a temperature 1000 K peaks in the

- (a) Visible range of frequencies                      (b) Infrared range of frequencies  
(c) Ultraviolet range of frequencies                      (d) Microwave range of frequencies

Ans: (a)

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Q5. The three principal moments of inertia of a methanol ( $\text{CH}_3\text{OH}$ ) molecule have the property  $I_x = I_y = I$  and  $I_z \neq I$ . The rotation energy eigenvalues are

- (a)  $\frac{\hbar^2}{2I}l(l+1) + \frac{\hbar^2 m_l^2}{2I} \left( \frac{1}{I_z} - \frac{1}{I} \right)$                       (b)  $\frac{\hbar^2}{2I}l(l+1)$   
 (c)  $\frac{\hbar^2 m_l^2}{2I} \left( \frac{1}{I_z} - \frac{1}{I} \right)$                       (d)  $\frac{\hbar^2}{2I}l(l+1) + \frac{\hbar^2 m_l^2}{2} \left( \frac{1}{I_z} + \frac{1}{I} \right)$

Ans: (a)

Solution:  $\text{CH}_3\text{OH}$  is example of symmetric rotar where  $I_x = I_y \neq I_z$  ( $I_x + I_y = I$  and  $I_z \neq I$ )

The classical expression for energy is  $E = \frac{1}{2I}(J_x^2 + J_y^2) + \frac{1}{2I_z}J_z^2$

This can be expressed in term of  $J^2 = J_x^2 + J_y^2 + J_z^2$  by adding and subtracting  $J_z^2$

$$E = \frac{1}{2I}J^2 + \left( \frac{1}{2I_z} - \frac{1}{2I} \right) J_z^2.$$

Quantum mechanically  $E = \frac{\hbar^2}{2I}J(J+1) + \frac{\hbar^2 m_J^2}{2} \left( \frac{1}{I_z} - \frac{1}{I} \right)$

Q6. Match the typical spectra of stable molecules with the corresponding wave-number range

- |                            |                                      |
|----------------------------|--------------------------------------|
| 1. Electronic spectra      | (i) $10^6 \text{ cm}^{-1}$ and above |
| 2. Rotational spectra      | (ii) $10^5 - 10^6 \text{ cm}^{-1}$   |
| 3. Molecule dissociation   | (iii) $10^8 - 10^2 \text{ cm}^{-1}$  |
| (a) 1 – ii, 2 – i, 3 – iii | (b) 1 – ii, 2 – iii, 3 – i           |
| (b) 1 – iii, 2 – ii, 3 – i | (d) 1 – i, 2 – ii, 3 – iii           |

Ans: (b)

Q7. Consider the operations  $P: \vec{r} \rightarrow -\vec{r}$  (parity) and  $T: t \rightarrow -t$  (time reversal). For the electric and magnetic fields  $\vec{E}$  and  $\vec{B}$ , which of the following set of transformations is correct?

- |   |   |
|---|---|
| (a) $P: \vec{E} \rightarrow -\vec{E}, \vec{B} \rightarrow \vec{B};$<br>$T: \vec{E} \rightarrow \vec{E}, \vec{B} \rightarrow -\vec{B}$ | (b) $P: \vec{E} \rightarrow \vec{E}, \vec{B} \rightarrow \vec{B};$<br>$T: \vec{E} \rightarrow \vec{E}, \vec{B} \rightarrow \vec{B}$ |
| (c) $P: \vec{E} \rightarrow -\vec{E}, \vec{B} \rightarrow \vec{B};$   | (d) $P: \vec{E} \rightarrow \vec{E}, \vec{B} \rightarrow -\vec{B};$   |

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$$T: \vec{E} \rightarrow -\vec{E}, \vec{B} \rightarrow -\vec{B}$$

$$T: \vec{E} \rightarrow -\vec{E}, \vec{B} \rightarrow \vec{B}$$

Ans: (b)

**Common Data Questions 8 and 9:**

In the presence of a weak magnetic field, atomic hydrogen undergoes the transition:

$${}^2P_{1/2} \rightarrow {}^1S_{1/2}, \text{ by emission of radiation}$$

Q8. The number of distinct spectral lines that are observed in the resultant Zeeman spectrum is

- (a) 2                                      (b) 3                                      (c) 4                                      (d) 6

Ans: (c)

Solution:  ${}^2P_{1/2} - {}^2S_{1/2}$  is sodium D1 lines and it has total 4 zeeman components.

Q9. The spectral line corresponding to the transition

$${}^2P_{1/2} \left( m_j = +\frac{1}{2} \right) \rightarrow {}^1S_{1/2} \left( m_j = -\frac{1}{2} \right)$$

is observed along the direction of the applied magnetic field. The emitted electromagnetic field is

- (a) Circularly polarized                      (b) Linearly polarized  
(c) Unpolarized                                  (d) Not emitted along the magnetic field direction

Ans: (a)

Solution: For  ${}^2P_{1/2} \left( m_j = +\frac{1}{2} \right) \rightarrow {}^1S_{1/2} \left( m_j = -\frac{1}{2} \right)$

Here  $\Delta m_j = +1$  gives  $\sigma^+$  component. In longitudinal observation  $\sigma^+$  is circularly polarized.

**GATE-2011**

Q10. The population inversion in a two layer material **CANNOT** be achieved by optical pumping because

- (a) the rate of upward transitions is equal to the rate of downward transitions  
(b) the upward transitions are forbidden but downward transitions are allowed  
(c) the upward transitions are allowed but downward transitions are forbidden

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(d) the spontaneous decay rate of the higher level is very low

Ans: (a)

Q11. A heavy symmetrical top is rotating about its own axis of symmetry (the  $z$ -axis). If  $I_1$ ,  $I_2$  and  $I_3$  are the principal moments of inertia along  $x$ ,  $y$  and  $z$  axes respectively, then

(a)  $I_2 = I_3; I_1 \neq I_2$       (b)  $I_1 = I_3; I_1 \neq I_2$       (c)  $I_1 = I_2; I_1 \neq I_3$       (d)  $I_1 \neq I_2 \neq I_3$

Ans: (c)

Q12. A neutron passing through a detector is detected because of

(a) the ionization it produces      (b) the scintillation light it produces

(c) the electron-hole pairs it produces

(d) the secondary particles produced in a nuclear reaction in the detector medium

Ans: (b)

Q13. An atom with one outer electron having orbital angular momentum  $l$  is placed in a weak magnetic field. The number of energy levels into which the higher total angular momentum state splits, is

(a)  $2l + 2$       (b)  $2l + 1$       (c)  $2l$       (d)  $2l - 1$

Ans: (b)

Q14. For a multi-electron atom  $l$ ,  $L$  and  $S$  specify the one-electron orbital angular momentum, total orbital angular momentum and total spin angular momentum, respectively. The selection rules for electric dipole transition between the two electronic energy levels, specified by  $l$ ,  $L$  and  $S$  are

(a)  $\Delta L = 0, \pm 1; \Delta S = 0; \Delta l = 0, \pm 1$       (b)  $\Delta L = 0, \pm 1; \Delta S = 0; \Delta l = \pm 1$

(c)  $\Delta L = 0, \pm 1; \Delta S = \pm 1; \Delta l = 0, \pm 1$       (d)  $\Delta L = 0, \pm 1; \Delta S = \pm 1; \Delta l = \pm 1$

Ans: (b)

Q15. The lifetime of an atomic state is 1 nanosecond. The natural line width of the spectral line in the emission spectrum of this state is of the order of

(a)  $10^{-10}$  eV      (b)  $10^{-9}$  eV      (c)  $10^{-6}$  eV      (d)  $10^{-4}$  eV

Ans: (c)

$$\text{Solution: } \Delta E = h\Delta\nu = \frac{h}{\Delta t} = \frac{6.625 \times 10^{-34} \text{ J} \cdot \text{s}}{10^{-9}} = \frac{6.625 \times 10^{-25}}{1.6 \times 10^{-19}} \text{ eV} = 4.14 \times 10^{-6} \text{ eV}$$

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- Q16. The degeneracy of an excited state of nitrogen atom having electronic configuration  $1s^2 2s^2 2p^2 3d^1$  is
- (a) 6                                      (b) 10                                      (c) 15                                      (d) 150

Ans: (b)

Solution: Degeneracy =  $2(2l + 1)$

- Q17. The far infrared rotational absorption spectrum of a diatomic molecule shows equilibrium lines with spacing  $20 \text{ cm}^{-1}$ . The position of the first Stokes line in the rotational Raman spectrum of this molecule is
- (a)  $20 \text{ cm}^{-1}$                               (b)  $40 \text{ cm}^{-1}$                               (c)  $60 \text{ cm}^{-1}$                               (d)  $120 \text{ cm}^{-1}$

Ans: (c)

Solution: Given  $2B = 20 \text{ cm}^{-1} \Rightarrow B = 10 \text{ cm}^{-1}$

The position of the first stokes line in the rotational Raman spectrum =  $6B$

$$\therefore \Delta \bar{\nu} = 6B = 6 \times 10 = 60 \text{ cm}^{-1}.$$

### GATE-2012

- Q18. The ground state of sodium atom ( $^{11}\text{Na}$ ) is a  $^2S_{1/2}$  state. The difference in energy levels arising in the presence of a weak external magnetic field  $B$ , given in terms of Bohr magneton,  $\mu_B$ , is
- (a)  $\mu_B B$                                       (b)  $2\mu_B B$                                       (c)  $4\mu_B B$                                       (d)  $6\mu_B B$

Ans: (b)

Solution: The energy separation in the Zeeman level is  $\Delta E = gM_J(\mu_B B)$

For  $^2S_{1/2}$  state;  $g = 2$  and  $M_J = \pm \frac{1}{2}$ . Therefore  $\Delta E_1 = +(\mu_B B)$  and  $\Delta E_2 = -(\mu_B B)$ .

$$\text{Thus } \Delta E = 2\mu_B B$$

- Q19. The first Stokes line of a rotational Raman spectrum is observed at  $12.96 \text{ cm}^{-1}$ . Considering the rigid rotor approximation, the rotational constant is given by
- (a)  $6.48 \text{ cm}^{-1}$                               (b)  $3.24 \text{ cm}^{-1}$                               (c)  $2.16 \text{ cm}^{-1}$                               (d)  $1.62 \text{ cm}^{-1}$

Ans: (c)

Solution: The first Stoke line of the Rotational Raman spectrum lies at  $= 6B$

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Thus  $6B = 12.96 \text{ cm}^{-1} \Rightarrow B = 1.16 \text{ cm}^{-1}$ .

Q20. Match the typical spectroscopic regions specified in **Group I** with the corresponding type of transitions in **Group II**.

**Group I**

(P) Infra-red region

(Q) Ultraviolet-visible region

(R) X-ray region

(S)  $\gamma$ -ray region

(a) (P, i); (Q, iii); (R, ii); (S, iv)

(c) (P, iii); (Q, i); (R, iv); (S, ii)

**Group II**

(i) electronic transitions involving valence electrons

(ii) nuclear transitions

(iii) vibrational transitions of molecules

(iv) transitions involving inner shell electrons

(b) (P, ii); (Q, iv); (R, i); (S, iii)

(d) (P, iv); (Q, i); (R, ii); (S, iii)

Ans: (c)

Q21. The term  $\{j_1, j_2\}_J$  arising from  $2s^1 3d^1$  electronic in  $j-j$  coupling scheme are

(a)  $\left\{\frac{1}{2}, \frac{3}{2}\right\}_{2,1}$  and  $\left\{\frac{1}{2}, \frac{5}{2}\right\}_{3,2}$

(b)  $\left\{\frac{1}{2}, \frac{1}{2}\right\}_{1,0}$  and  $\left\{\frac{1}{2}, \frac{3}{2}\right\}_{2,1}$

(c)  $\left\{\frac{1}{2}, \frac{1}{2}\right\}_{1,0}$  and  $\left\{\frac{1}{2}, \frac{5}{2}\right\}_{3,2}$

(d)  $\left\{\frac{3}{2}, \frac{1}{2}\right\}_{2,1}$  and  $\left\{\frac{1}{2}, \frac{5}{2}\right\}_{3,2}$

Ans: (c)

Q22. The equilibrium vibration frequency for an oscillator is observed at  $2990 \text{ cm}^{-1}$ . The ratio of the frequencies corresponding to the first and the fundamental spectral lines is 1.96.

Considering the oscillator to be anharmonic, the anharmonicity constant is

(a) 0.005

(b) 0.02

(c) 0.05

(d) 0.1

Ans: (b)

Solution:  $\omega_e(1 - 2x_e) = 2990 \text{ cm}^{-1}$  and  $\frac{2\omega_e(1 - 3x_e)}{\omega_e(1 - 2x_e)} = 1.96 \text{ cm}^{-1} \Rightarrow \frac{(1 - 3x_e)}{(1 - 2x_e)} = 0.98 \text{ cm}^{-1}$

$\Rightarrow x_e = 0.02$ .

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**GATE-2013**

Q23. The number of spectral lines allowed in the spectrum for the  $3^2D \rightarrow 3^2P$  transition in sodium is \_\_\_\_\_.

Ans: 28

Solution: The numbers of Zeeman components for  $^2D_{5/2} \rightarrow ^2P_{3/2}$  transition = 12

The numbers of Zeeman components for  $^2D_{3/2} \rightarrow ^2P_{3/2}$  transition = 10

The numbers of Zeeman components for  $^2D_{3/2} \rightarrow ^2P_{1/2}$  transition = 6

Q24. In a normal Zeeman Effect experiment, spectral splitting of the line at the wavelength 643.8 nm corresponding to the transition  $5^1D_2 \rightarrow 5^1P_1$  of cadmium atoms is to be observed. The spectrometer has a resolution of 0.01 nm. Minimum magnetic field needed to observe this is ( $m_e = 9.1 \times 10^{-31} \text{ kg}$ ,  $e = 1.6 \times 10^{-19} \text{ C}$ ,  $c = 3 \times 10^8 \text{ m/s}$ )

- (a) 0.26T                      (b) 0.52T                      (c) 2.6T                      (d) 5.2T

Ans: (b)

Solution: Separation of Zeeman Components

$$\Delta \nu = \frac{eB}{4\pi m} \Rightarrow \Delta \lambda = \frac{\lambda^2}{c} \Delta \nu = \frac{\lambda^2}{c} \frac{eB}{4\pi m}$$

$$B = \frac{4\pi m c}{e} \frac{\Delta \lambda}{\lambda^2} = \frac{4 \times 3.14 \times 9.1 \times 10^{-31} \times 3 \times 10^8}{1.6 \times 10^{-19}} \frac{0.01 \times 10^{-9}}{(643.8 \times 10^{-9})^2} = 0.514T$$

Q25. The spacing between vibrational energy levels in CO molecule is found to be  $8.44 \times 10^{-2} \text{ eV}$ . Given that the reduced mass of CO is  $1.14 \times 10^{-26} \text{ kg}$ , Planck's constant is  $6.626 \times 10^{-34} \text{ Js}$  and  $1 \text{ eV} = 1.6 \times 10^{-19} \text{ J}$ . The force constant of the bond in CO molecule is

- (a) 1.87 N/m                      (b) 18.7 N/m                      (c) 187 N/m                      (d) 1870 N/m

Ans: (c)

Solution: The energy of the quantum harmonic oscillator is

$$E = h\nu \left( n + \frac{1}{2} \right), \quad n = 0, 1, 2, \dots$$

$$\text{The frequency of oscillation is } \nu = \frac{1}{2\pi} \sqrt{\frac{k}{\mu}}$$

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Where  $k$  = Spring constant and  $\mu$  = reduced mass

The energy levels are equally spaced with energy separation of  $\Delta E = h\nu = \frac{h}{2\pi} \sqrt{\frac{k}{\mu}}$

$$k = \left( \frac{2\pi}{h} \Delta E \right)^2 \mu = \left( \frac{2 \times 3.14}{6.626 \times 10^{-34}} \times 8.44 \times 10^{-2} \times 1.6 \times 10^{-19} \right)^2 \times 1.14 \times 10^{-26} = 186.7 \text{ N/m}$$

### GATE-2014

Q26. The number of normal Zeeman splitting components of  $^1P \rightarrow ^1D$  transition is

- (a) 3                                      (b) 4                                      (c) 8                                      (d) 9

Ans: (a)

Solution: This is singlet transition.

Q27. The moment of inertia of a rigid diatomic molecule  $A$  is 6 times that of another rigid diatomic molecule  $B$ . If the rotational energies of the two molecules are equal, then the corresponding values of the rotational quantum numbers  $J_A$  and  $J_B$  are

- (a)  $J_A = 2, J_B = 1$                                       (b)  $J_A = 3, J_B = 1$   
 (c)  $J_A = 5, J_B = 0$                                       (d)  $J_A = 6, J_B = 1$

Ans: (b)

Solution:  $\frac{J_A(J_A + 1)}{J_B(J_B + 1)} = \frac{I_B}{I_A} \Rightarrow \frac{I_B}{6I_B}$

$$J_A = 6, J_B = 1$$

Q28. The value of the magnetic field required to maintain non-relativistic protons of energy  $1\text{MeV}$  in a circular orbit of radius  $100\text{mm}$  is \_\_\_\_\_ Tesla

(Given:  $m_p = 1.67 \times 10^{-27} \text{ kg}$ ,  $e = 1.6 \times 10^{-19} \text{ C}$ )

Ans: 1.44

Solution:  $\frac{mv^2}{r} = qvB$ ,  $E = \frac{1}{2}mv^2 \Rightarrow B = \frac{\sqrt{2mE}}{qr} = 1.44$

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Q29. Neutrons moving with speed  $10^3 \text{ m/s}$  are used for the determination of crystal structure.

If the Bragg angle for the first order diffraction is  $30^\circ$  the interplanar spacing of the

crystal is \_\_\_\_\_  $\text{\AA}$

(Given:  $m_n = 1.675 \times 10^{-27} \text{ kg}$   $h = 6.626 \times 10^{-34} \text{ J.s}$ )

Ans: 4

Solution:  $2d \sin \theta = \lambda = \frac{h}{mv} \therefore d = 4 \text{\AA}$

Q30. The emission wavelength for the transition  $D_2 \rightarrow F_3$  is  $3122 \text{ \AA}$ . The ratio of population of the final to the initial states at a temperature  $5000 \text{ K}$  is

( $h = 6.626 \times 10^{-34} \text{ J.s}$ ,  $c = 3 \times 10^8 \text{ m/s}$   $k_B = 1.380 \times 10^{-23} \text{ J/K}$ )

(a)  $2.03 \times 10^{-5}$  (b)  $4.02 \times 10^{-5}$  (c)  $7.02 \times 10^{-5}$  (d)  $9.83 \times 10^{-5}$

Ans: (c)

Solution:  $\frac{N_f}{N_i} = \frac{(2J_f + 1)}{(2J_i + 1)} e^{-\frac{hc}{\lambda k_B T}} = \frac{5}{7} e^{-9.227641144} = 7.02 \times 10^{-5}$

### GATE-2015

Q31. In a rigid rotator of mass  $M$ , if the energy of the first excited state is  $(1 \text{ meV})$ , then the fourth excited state energy (in  $\text{meV}$ ) is \_\_\_\_\_.

Ans.: 6

Solution:  $\therefore E \propto J(J+1)$  where  $J = 0, 1, 2, 3..$

$\Rightarrow \frac{E_4}{E_1} = \frac{3(3+1)}{1(1+1)} \Rightarrow E_4 = 6E_1 = 6 \text{ meV}$  where  $J = 0, 1, 2, 3..$

Q32. The binding energy per molecule of  $\text{NaCl}$  (lattice parameter is  $0.563 \text{ nm}$ ) is  $7.956 \text{ eV}$ .

The repulsive term of the potential is of the form  $\frac{K}{r^9}$ , where  $K$  is a constant. The value

of the Madelung constant is \_\_\_\_\_ (upto three decimal places)

(Electron charge  $e = -1.6 \times 10^{-19} \text{ C}$ ;  $\epsilon_0 = 8.854 \times 10^{-12} \text{ C}^2 \text{ N}^{-1} \text{ m}^{-2}$ )

Ans.: 1.68

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Solution: The total energy of one ion due to the presence of all others in NaCl crystal is (considering univalent ions)

$$U(r) = -\frac{Ae^2}{4\pi\epsilon_0 r} + \frac{K}{r^n} \text{ where } A \text{ is Modelung Constant.}$$

The potential energy will be minimum at the equilibrium spacing  $r_0$ .

$$\text{Thus } \left[ \frac{dU}{dr} \right]_{r=r_0} = \left[ \frac{Ae^2}{4\pi\epsilon_0 r_0^2} - \frac{Kn}{r_0^{n+1}} \right] = 0 \Rightarrow K = \frac{Ae^2 r_0^{n-1}}{4\pi\epsilon_0 n}$$

Thus, Binding energy of molecule or lattice energy is

$$U_0 = [U]_{r=r_0} = -\frac{Ae^2}{4\pi\epsilon_0 r_0} + \frac{Ae^2 r_0^{n-1}}{4\pi\epsilon_0 n r_0^n} = -\left[ \frac{Ae^2}{4\pi\epsilon_0 r_0} \right] \left[ \frac{n-1}{n} \right]$$

Given repulsive term of the potential is  $\frac{K}{r^9} (\because n=9)$

Also binding energy per molecule is  $U_0 = 7.95 \text{ eV}$

The Modelung constant is  $A = U_0 \times \frac{4\pi\epsilon_0 r_0}{e^2} \times \frac{n}{n-1}$

Given, the lattice parameter is  $a = 0.563 \text{ nm}$ , thus, the interatomic separation is

$$r_0 = \frac{a}{2} = 0.282 \text{ nm}$$

$$\Rightarrow A = 7.95 \times 1.67 \times 10^{-19} \text{ J} \times \frac{4 \times 3.14 \times (8.85 \times 10^{-12} \text{ C}^2 \text{ N}^{-1} \text{ m}^{-2}) \times (0.282 \times 10^{-9})^9}{(1.67 \times 10^{-19} \text{ J})^2} \times \frac{9}{8}$$

$$\Rightarrow A = \frac{7.95 \times 4 \times 3.14 \times 8.85 \times 0.282 \times 9}{1.67 \times 8} \times 10^{-2} \Rightarrow A = 1.68$$

Q33. Match the phrases in Group I and Group II and identify the correct option.

Group I

- (P) Electron spin resonance (ESR)
- (Q) Nuclear magnetic resonance (NMR)
- (R) Transition between vibrational states of a molecule
- (S) Electronic transition

Group II

- (i) radio frequency
- (ii) visible range frequency
- (iii) microwave frequency
- (iv) far-infrared range

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(a) (P-i), (Q-ii), (R-iii), (S-iv)

(b) (P-ii), (Q-i), (R-iv), (S-iii)

(c) (P-iii), (Q-iv), (R-i), (S-ii)

(d) (P-iii), (Q-i), (R-iv), (S-ii)

Ans.: (d)

Solution: (P) Electron spin resonance (ESR) is achieved by Microwave frequency (iii)

(Q): Nuclear magnetic resonance (NMR) is achieved by Radio frequency (i)

(R): Transition between vibrational states of a molecule is achieved by radiation of far infrared range (iv)

(S): Electronic transition is achieved by visible radiation (ii)

Q34. The excitation wavelength of laser in a Raman effect experiment is  $546\text{ nm}$ . If the Stokes' line is observed at  $552\text{ nm}$ , then the wavenumber of the anti-Stokes' line (in  $\text{cm}^{-1}$ ) is \_\_\_\_\_

Ans.: 18514

Solution: Raman displacement is

$$\Delta \bar{\nu} = \bar{\nu}_{AS} - \bar{\nu}_0 = \bar{\nu}_0 - \bar{\nu}_S \quad \text{or} \quad \Delta \bar{\nu} = \frac{1}{\lambda_{AS}} - \frac{1}{\lambda_0} = \frac{1}{\lambda_0} - \frac{1}{\lambda_S}$$

where  $\lambda_{AS}$ ,  $\lambda_0$ ,  $\lambda_S$  are wavelength of anti-stoke, exciting & stoke line.

From above relation we can write

$$\frac{1}{\lambda_{AS}} - \frac{1}{\lambda_0} = \frac{1}{\lambda_0} - \frac{1}{\lambda_S} \Rightarrow \frac{1}{\lambda_{AS}} = \frac{2}{\lambda_0} - \frac{1}{\lambda_S} \Rightarrow \frac{1}{\lambda_{AS}} = \frac{2\lambda_S - \lambda_0}{\lambda_0\lambda_S} \Rightarrow \lambda_{AS} = \frac{\lambda_0\lambda_S}{2\lambda_S - \lambda_0}$$

$$\therefore \lambda_{AS} = \frac{(546 \times 10^{-9} \text{ m})(552 \times 10^{-9} \text{ m})}{(2 \times 552 \times 10^{-9} \text{ m} - 546 \times 10^{-9} \text{ m})} = \frac{546 \times 552}{558} \times 10^{-9} \text{ m}$$

$$\Rightarrow \lambda_{AS} = 540.129 \times 10^{-9} \text{ m} = 540.129 \times 10^{-7} \text{ cm}$$

$$\text{Anti-stoke wavenumber is } \bar{\nu}_{AS} = \frac{1}{\lambda_{AS}} = \frac{1}{540.129 \times 10^{-7} \text{ cm}} = 18514 \text{ cm}^{-1}$$

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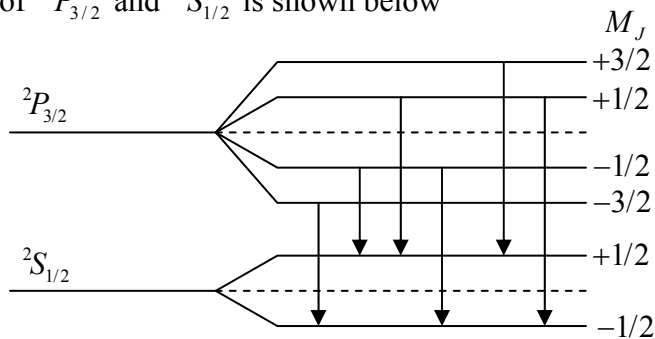
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Q35. The number of permitted transitions from  ${}^2P_{3/2} \rightarrow {}^2S_{1/2}$  in the presence of a weak magnetic field is \_\_\_\_\_

Ans. : 6

Solution: Zeeman splitting of  ${}^2P_{3/2}$  and  ${}^2S_{1/2}$  is shown below



The selection rule for Zeeman transactions are

$$\Delta M_j = 0, \pm 1 \quad (0 \rightarrow 0 \Delta \text{ if } J = 0)$$

There are total six transition in accordance with above selection rules.

### GATE-2016

Q36. The molecule  ${}^{17}O_2$

- (a) Raman active but not NMR (nuclear magnetic resonance) active.
- (b) Infrared active and Raman active but not NMR active.
- (c) Raman active and NMR active.
- (d) Only NMR active.

Ans.: (c)

Solution: (i) Molecule  ${}^{17}O_2$  can not absorb infrared as there is no change in dipole moment during vibration. Thus  ${}^{17}O_2$  is infrared inactive.

(ii) Molecule  ${}^{17}O_2$  shows change in polarizability during rotation. Thus it is Raman active molecule.

(iii) The nucleus of  ${}^{17}O$  has spin  $\frac{5}{2}$ , therefore it is NMR active.

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Q37. There are four electrons in the  $3d$  shell of an isolated atom. The total magnetic moment of the atom in units of Bohr magneton is \_\_\_\_\_.

Ans.: 0

Solution: The configuration leads to  $S = 2$  and  $L = 2$

Since it is the case of less than half filled sub shell, thus according to Hund's rules, lower  $J$  will be in ground state.

$$M_L = -2 \quad -1 \quad 0 \quad +1 \quad +2$$

$$\therefore J = L - S = 0. \text{ Thus } \bar{\mu} = 0$$

$$3d^4 = \begin{array}{|c|c|c|c|c|} \hline \uparrow & \uparrow & \uparrow & \uparrow & \square \\ \hline \end{array}$$

Q38. Which of the following transitions is NOT allowed in the case of an atom, according to the electric dipole radiation selection rule?

- (a)  $2s - 1s$                       (b)  $2p - 1s$                       (c)  $2p - 2s$                       (d)  $3d - 2p$

Ans.: (a)

Solution: In electric dipole transition,  $\Delta l = \pm 1$ . Thus in transition  $2s \rightarrow 1s$ ,  $\Delta l = 0$ . It violates the selection rule and hence not allowed.

Q39. The number of spectroscopic terms resulting from the  $L.S$  coupling of a  $3p$  electron and a  $3d$  electron is \_\_\_\_\_.

Ans.: 12

Solution: For  $3p^1 3d^1$ :  $s_1 = \frac{1}{2}, s_2 = \frac{1}{2} \therefore S = 0, 1$

$$l_1 = 1, l_2 = 2 \quad \therefore L = 1, 2, 3$$

$$S = 0, L = 1 \Rightarrow J = 1 \quad \therefore \text{Term} = {}^1P_1$$

$$S = 0, L = 2 \Rightarrow J = 2 \quad \therefore \text{Term} = {}^1D_2$$

$$S = 0, L = 3 \Rightarrow J = 3 \quad \therefore \text{Term} = {}^1F_3$$

$$S = 1, L = 1 \Rightarrow J = 0, 1, 2 \quad \therefore \text{Terms} = {}^3P_0, {}^3P_1, {}^3P_2$$

$$S = 1, L = 2 \Rightarrow J = 1, 2, 3 \quad \therefore \text{Terms} = {}^3D_1, {}^3D_2, {}^3D_3$$

$$S = 1, L = 3 \Rightarrow J = 2, 3, 4 \quad \therefore \text{Terms} = {}^3F_2, {}^3F_3, {}^3F_4$$

Thus total number of spectroscopic terms are 12.

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